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Chapter 3 MCQs : Atoms and Molecules

Q.1. Law of conservation of mass was for the first time proposed by-

(a) Lavoiser (b) Richter
(c) Proust (d) Dalton

Q.2. Law of constant proportion was for the first time proposed by-

(a) Lavoiser (b) Richter
(c) Proust (d) Dalton

Q.3. The symbol of copper is-

(a) C (b) Cu
(c) Co (d) Ce

Q.4. The symbol of potassium is-

(a) P (b) Po
(c) K (d) Pt

Q.5. The symbol of which of the following elements has been derived from its Latin name-

(a) P (b) K
(c) Mg (d) S

Q.6. The symbol of which of the following elements has not been derived from its Latin name-

(a) Na (b) Cu
(c) Mg (d) Fe

Q.7. The symbol of which of the following elements is based on its colour ?

(a) Na (b) Cu
(c) Mg (d) Au

Q.8. The atomic mass of chlorine is-

(a) 23 u (b) 31 u
(c) 35.5 u (d) None of the above

Q.9. The atomic mass of iron is-

(a) 23 u (b) 31 u
(c) 35.5 u (d) 56 u

Q.10. The smallest particle of an element which takes part in a chemical reaction without being decomposed is called-

(a) Atom (b) Molecule
(c) Ions (d) None of the above

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Q.11. When two or more elements combine together in a fixed ratio then they form-

(a) Compound (b) Mixture
(c) Ions (d) Homogeneous mixture

Q.12. The formula of aluminium oxide is-

(a) Al_2O (b) Al_2O_3
(c) Al_3O_4 (d) Al_3O

Q.13. The formula of calcium phosphate is-

(a) $Ca_3(PO_4)_2$ (b) $Ca_2(PO_4)_3$
(c) Ca_2PO_4 (d) $Ca(PO_4)_2$

Q.14. The formula mass of NaCl is-

(a) 23 u (b) 35.5 u
(c) 58.5 u (d) None of the above

Q.15. The formula mass of $CaCO_3$ is-

(a) 23 u (b) 35.5 u
(c) 58.5 u (d) 100 u

Q.16. The molecular mass of H_2SO_4 is-

(a) 98 u (b) 92 u
(c) 94 u (d) 96 u

Q.17. The molecular mass of CO_2 is-

(a) 24 u (b) 32 u
(c) 44 u (d) None of the above

Q.18. The atomicity of H_2SO_4 is-

(a) 7 (b) 3
(c) 2 (d) 5

Q.19. Which of the following has the highest atomicity ?

(a) Phosphorus (b) Ozone
(c) Sulphur (d) Chlorine

Q.20. Which of the following is universally accepted standard for showing relative atomic mass of different elements ?

(a) H (b) O
(c) C (d) C-12 isotope

Q.21. Which of the following corresponds to 1 u ?

(a) Molecular mass of H_2 (b) Atomic mass of oxygen
(c) Atomic mass of C-12 isotope (d) $\frac{1}{12}\times$ Mass of one C-12 atom

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Q.22. 1 nanometer is equal to-

(a) 10^{-3} m (b) 10^{-6} m

(c) 10^{-9} m (d) 10^{-12} m

Q.23. Which of the following is called 1 mole-

(a) 1 GAM of an element (b) 1 GMM of a substance

(c) Both the above (d) None of the above

Q.24. Which of the following corresponds to 1 u-

(a) Mass of one atom of H (b) $\frac{1}{12}\times$ Mass of one atom of C-12

(c) $\frac{1}{16}\times$ Mass of one atom of O-16 (d) All the above

Q.25. The molecular mass of hydrogen is-

(a) 1 g (b) 1 u

(c) 2 g (d) 2 u

ANSWERS

1. (a)	**2. (c)**	**3. (b)**	**4. (c)**	**5. (b)**
6. (c)	**7. (d)**	**8. (c)**	**9. (d)**	**10. (a)**
11. (a)	**12. (b)**	**13. (a)**	**14. (c)**	**15. (d)**
16. (a)	**17. (c)**	**18. (a)**	**19. (c)**	**20. (d)**
21. (d)	**22. (c)**	**23. (c)**	**24. (d)**	**25. (d)**

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