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# OBJECTIVE TEST

## Topic – Periodic Classification

**Time – 1 Hr.** **M.M.- 50**

**Q.1. Which of the following is considered to be fundamental properties of the elements-**

**(a) Atomic mass** **(b) Atomic number**
**(c) Both** **(d) Mass number**

**Q.2. Total number of series in Mendeleev's Original Periodic Table were -**

**(a) 7** **(b) 10** **(c) 18** **(d) 8**

**Q.3. Mendeleev's Modern Periodic Table is based on-**

**(a) Atomic mass** **(b) Atomic number**
**(c) Both** **(d) Mass number**

**Q.4. Long Form of Periodic Table is based upon-**

**(a) Atomic mass** **(b) Atomic number**
**(c) Both** **(d) Electronic configuration**

**Q.5.Which of the following is not the example of inverted pair-**

**(a) Ar & K** **(b) Co & Ni**
**(c) Cu & Zn** **(d) Te & I**

**Q.6. Representative elements include following blocks of the Periodic Table-**

**(a) s-Block** **(b) p-Block**
**(c) d- Block** **(d) Both s & p-Block**

**Q.7. Transition elements include following blocks of the Periodic Table-**

**(a) s-Block** **(b) p-Block**
**(c) d- Block** **(d) Both d & f-Block**

**Q.8. Lanthenides (4f-series) belong to the period -**

**(a) IV** **(b) V** **(c) VI** **(d) VII**

**Q.9. Actinides (5f-series) belong to the period -**

**(a) IV** **(b) V** **(c) VI** **(d) VII**

**Q.10. The outer shell electronic configuration of p-Block elements is-**

**(a) $ns^2$, $np^{1-6}$** **(b) $ns^2$, $(n-1)p^{1-6}$**
**(c) $ns^2$, $(n-1)d^{1-10}$** **(d) $ns^2$, $(n-1)d^{0-1}$,$(n-2)f^{1-14}$**

**Q.11. The outer shell electronic configuration of d-Block elements is-**

**(a) $ns^2$, $np^{1-6}$** **(b) $ns^2$, $(n-1)p^{1-6}$**
**(c) $ns^2$, $(n-1)d^{1-10}$** **(d) $ns^2$, $(n-1)d^{0-1}$,$(n-2)f^{1-14}$**

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**Q.12.Which of the following has biggest atomic size-**

**(a) Li (b) Be (c) O (d) F**

**Q.13. Which of the following has biggest atomic size-**

**(a) Na (b) Ne (c) S (d) Cl**

**Q.14.The correct sequence of atomic size is-**

**(a) Na > Mg > Al > Si (b) Si > Al > Mg > Na**

**(c) Na > Al > Mg > Si (d) Na > Si > Al > Mg**

**Q.15. The correct sequence of atomic size is-**

**(a) Ne > Li > O > F (b) Li > Ne > O > F**

**(c) Ne > F > O > Li (d) Li > O > F > Ne**

**Q.16. The correct sequence of atomic size is-**

**(a) Na > K > O > S (b) S > O > Na > K**

**(c) K > Na > S > O (d) K > S > Na > O**

**Q.17. The correct sequence of atomic size is-**

**(a) Li > Na > O > S (b) S > Na > O > Li**

**(c) Na > Li > S > O (d) Na > S > Li > O**

**Q.18. The correct sequence of atomic size is-**

**(a) P > N > O > S (b) P > S > O > N**

**(c) P > S > N > O (d) S > P > O > N**

**Q.19. Which of the following has biggest Ionic size-**

**(a) $Na^+$ (b) $N^{3-}$ (c) $O^{2-}$ (d) $F^-$**

**Q.20. Which of the following has smallest Ionic size-**

**(a) $Na^+$ (b) $N^{3-}$ (c) $O^{2-}$ (d) $F^-$**

**Q.21. The correct sequence of Ionic size is-**

**(a) Ne > $Na^+$ > $Mg^{2+}$ > $Al^{3+}$ (b) $Na^+$ > Ne > $Mg^{2+}$ > $Al^{3+}$**

**(c) $Al^{3+}$ > $Mg^{2+}$ > $Na^+$ > Ne (d) Unpredictable**

**Q.22. The correct sequence of Ionic size is-**

**(a) $N^{3-}$ > $O^{2-}$ > $F^-$ > $Na^+$ (b) $Na^+$ > $F^-$ > $O^{2-}$ > $N^{3-}$**

**(c) $Na^+$ > $O^{2-}$ > $F^-$ > $N^{3-}$ (d) Unpredictable**

**Q.23.Which of the following has the highest ionization enthalpy-**

**(a) Ne (b) Na (c) Ne (d) He**

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**Q.24. Which of the following has the least ionization enthalpy-**

**(a) Na (b) Cs (c) Mg (d) Xe**

**Q.25. Which of the following has the most negative electron gain enthalpy-**

**(a) N (b) O (c) F (d) Cl**

**Q.26. Which of the following has highest electronegativity-**

**(a) N (b) O (c) F (d) Cl**

**Q.27. The correct sequence of ionization enthalpy is-**

**(a) C > N > O > F (b) F > N > O > C**

**(c) F > O > N > C (d) N > O > C > F**

**Q.28. The correct sequence of ionization enthalpy is-**

**(a) P > N > O > S (b) O > N > S > P**

**(c) P > S > N > O (d) N > O > P > S**

**Q.29. The correct sequence of ionization enthalpy is-**

**(a) Li > Be > B > C (b) C > Be > B > Li**

**(c) C > B > Be > Li (d) None**

**Q.30. The correct sequence of ionization enthalpy is-**

**(a) C > N > O > F (b) F > O > N > C**

**(c) F > N > O > C (d) None**

**Q.31. The correct sequence of negative value of electron gain enthalpy is-**

**(a) P > N > O > S (b) O > N > S > P**

**(c) S > P > O > N (d) None**

**Q.32. The correct sequence of negative value of electron gain enthalpy is-**

**(a) Cl > F > O > S (b) O > F > S > Cl**

**(c) Cl > S > S > O (d) None**

**Q.33. An element X having its atomic number 50 would belong to-**

**(a) $V^{th}$ Period & $14^{th}$ Group (b) $IV^{th}$ Period & $14^{th}$ Group**

**(c) $VII^{th}$ Period & $14^{th}$ Group (d) None**

**Q.34. Which of the following is a lanthenide –**

**(a) La (b) Ce (c) Fr (d) Te**

**Q.35.Which of the following is not a lanthenide -**

**(a) La (b) Ce (c) Lu (d) Pm**

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**Q.36. Which of the following is an actinide –**

**(b) Ac (b) Th (c) Os (d) Ra**

**Q.37.Which of the following is not an actinide -**

**(b)Ac (b) Am (c) U (d) Np**

**Q.38.In which of the following pairs both the elements have comparable atomic size -**

**(a) Zr & Hf (b) Nb & Ta (c) Mo & W (d) All**

**Q.39.Which of the following is not the property of d-block elements -**

**(a) Variable oxidation states (b) Flame colouration**
**(c) Complex ion formation (d) Alloy formation**

**Q.40.Which of the following is a peculiar property of s-block elements -**

**(a) Variable oxidation states (b) Flame colouration**
**(c) Complex ion formation (d) Coloured ion formation**

**Q.41. Which of the following has the highest metallic nature-**

**(a) Cl (b) Br (c) I (d) F**

**Q.42.Which of the following does not impart colour to the Bunsen flame –**

**(a) Be (b) Mg (c) Both (d) Ca**

**Q.43.Unnilpentium is the IUPAC name of a newly discovered element. The atomic number of this element will be -**

**(a) 105 (b) 115 (c) 155 (d) None**

**Q.44. Which of the following orbitals has the highest penetration power-**

**(a) s (b) p (c) d (d) f**

**Q.45. Which of the following orbitals shows the highest screening effect-**

**(a) s (b) p (c) d (d) f**

**Q.46. Which of the following is an amphoteric oxide-**

**(a) $Na_2O$ (b) MgO**
**(c) $Al_2O_3$ (d) $SiO_2$**

**Q.47. Which of the following is an acidic oxide-**

**(a) $Na_2O$ (b) MgO**
**(c) $Al_2O_3$ (d) $Mn_2O_7$**

**Q.48.Which of the following is the most acidic behaviour-**

**(a) HClO (b) $HClO_2$ (c) $HClO_3$ (d) $HClO_4$**

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**Q.49.Which of the following shows the most oxidizing behaviour-**

**(a) $HClO$ (b) $HClO_2$ (c) $HClO_3$ (d) $HClO_4$**

**Q.50.Which of the following is thermally most stable-**

**(a) $BeCO_3$ (b) $MgCO_3$ (c) $CaCO_3$ (d) $BaCO_3$**

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